

<u>Cation Formula</u>	<u>Cation Name</u>	<u>Anion Formula</u>	<u>Anion Name</u>
H ⁺	Hydrogen	BO ₃ ⁻³	Borate
Li ⁺	Lithium	CO ₃ ⁻²	Carbonate
Na ⁺	Sodium	HCO ₃ ⁻	Bicarbonate
K ⁺	Potassium	NO ₂ ⁻	Nitrite
Be ⁺²	Beryllium	NO ₃ ⁻	Nitrate
Mg ⁺²	Magnesium	PO ₂ ⁻³	Hypophosphite
Ca ⁺²	Calcium	PO ₃ ⁻³	Phosphite
Sr ⁺²	Strontium	PO ₄ ⁻³	Phosphate
Ba ⁺²	Barium	HPO ₄ ⁻²	Biphosphate
B ⁺³	Boron	H ₂ PO ₄ ⁻	Hydrogen Biphosphate
Al ⁺³	Aluminum	S ⁻²	Sulfide
Zn ⁺²	Zinc	HS ⁻	Bisulfide
Cd ⁺²	Cadmium	SO ₃ ⁻²	Sulfite
Hg ⁺²	Mercuric (Mercury II)	SO ₄ ⁻²	Sulfate
Hg ₂ ⁺²	Mercurous (Mercury I)	HSO ₄ ⁻	Bisulfate
Cu ⁺²	Cupric (Copper II)	F ⁻	Fluoride
Cu ⁺	Cuprous (Copper I)	FO ⁻	Hypofluorite
Ni ⁺³	Nickelic (Nickel III)	FO ₂ ⁻	Fluorite
Ni ⁺²	Nickelous (Nickel II)	FO ₃ ⁻	Fluorate
Fe ⁺³	Ferric (Iron III)	FO ₄ ⁻	Perfluorate
Fe ⁺²	Ferrous (Iron II)	Cl ⁻	Chloride
Ag ⁺	Silver	ClO ⁻	Hypochlorite
NH ₄ ⁺	Ammonium	ClO ₂ ⁻	Chlorite
Sn ⁺⁴	Stannic (Tin IV)	ClO ₃ ⁻	Chlorate
Sn ⁺²	Stannous (Tin II)	ClO ₄ ⁻	Perchlorate
Pb ⁺⁴	Plumbic (Lead IV)	Br ⁻	Bromide
Pb ⁺²	Plumbous (Lead II)	BrO ⁻	Hypobromite
Mn ⁺²	Manganous (Mn II)	BrO ₂ ⁻	Bromite
Mn ⁺³	Manganic (Mn III)	BrO ₃ ⁻	Bromate
Mn ⁺⁴	Manganite (Mn IV)	BrO ₄ ⁻	Perbromate
Cr ⁺²	Chromous (Chromium II)	I ⁻	Iodide
Cr ⁺³	Chromic (Chromium III)	IO ⁻	Hypoiodite
Co ⁺²	Cobaltous (Cobalt II)	IO ₂ ⁻	Iodite
Co ⁺³	Cobaltic (Cobalt III)	IO ₃ ⁻	Iodate
Sb ⁺³	Antimonous (Sb III)	IO ₆ ⁻⁵	Periodate
Sb ⁺⁵	Antimonic (Sb V)	OH ⁻	Hydroxide
PH ₄ ⁺	Phosphonium	O ⁻²	Oxide
H ₃ O ⁺	Hydronium	As ⁻²	Arsenide
H ₃ S ⁺	Sulfonium	AsO ₄ ⁻³	Arsenate
		AsS ₄ ⁻³	Thioarsenate
		S ₂ O ₃ ⁻²	Thiosulfate
		Se ⁻²	Selenide
		CrO ₄ ⁻²	Chromate
		Cr ₂ O ₇ ⁻²	Dichromate
		MnO ₄ ⁻	Permanganate (Mn VII)
		MnO ₄ ⁻²	Manganate (Mn VI)
		CH ₃ CO ₂ ⁻	Acetate
		CN ⁻	Cyanide
		OCN ⁻	Cyanate
		CNO ⁻	Fulminate
		SCN ⁻	Thiocyanate
		Fe(CN) ₆ ⁻⁴	Ferrocyanide
		Fe(CN) ₆ ⁻³	Ferricyanide
		N ₃ ⁻	Azide
		N ⁻³	Nitride
		P ⁻³	Phosphide